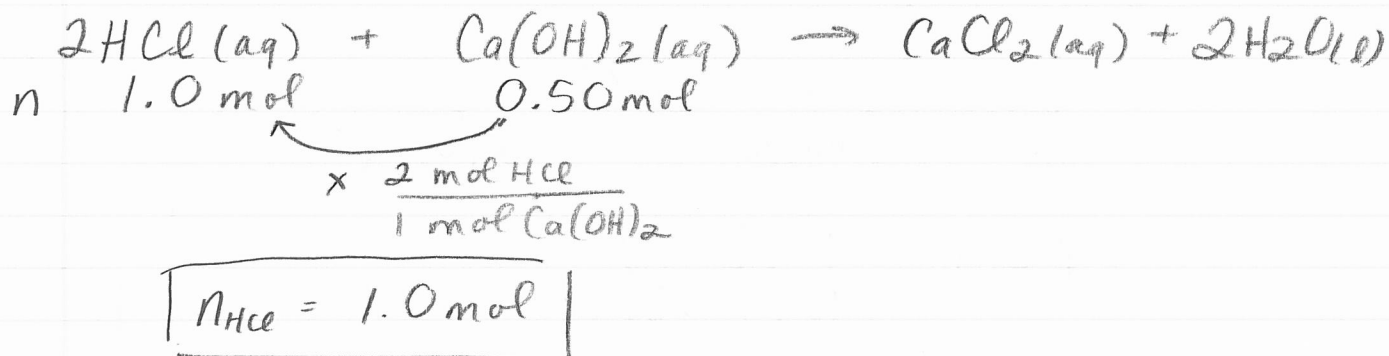
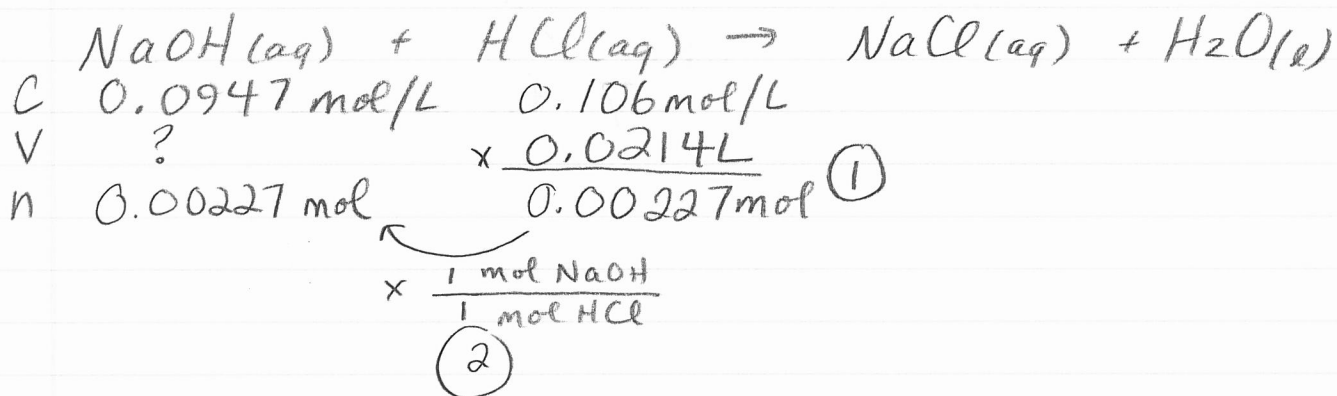


Example 1: Neutralization

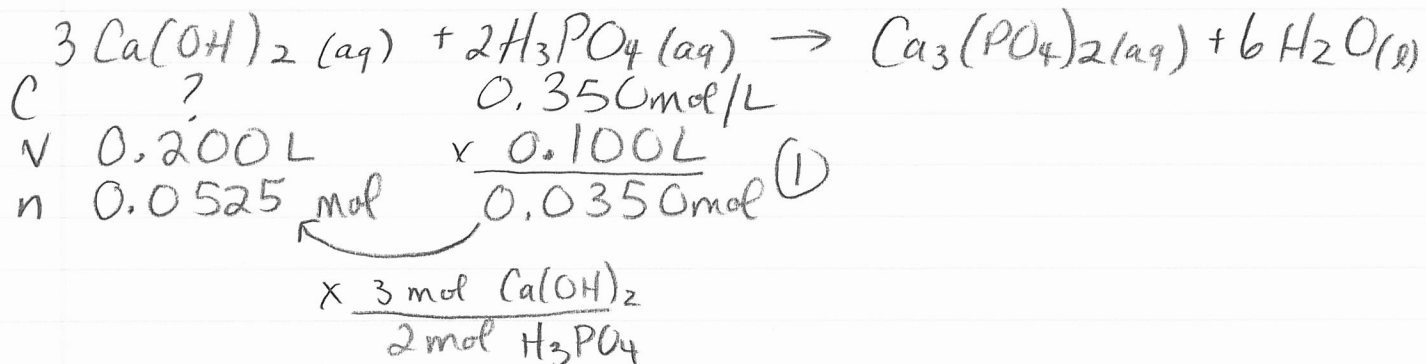


Example 2: Neutralization



$$\textcircled{3} \quad V_{\text{NaOH}} = \frac{n}{c} = \frac{0.00227}{0.0947} = \boxed{0.0240 \text{ L}}$$

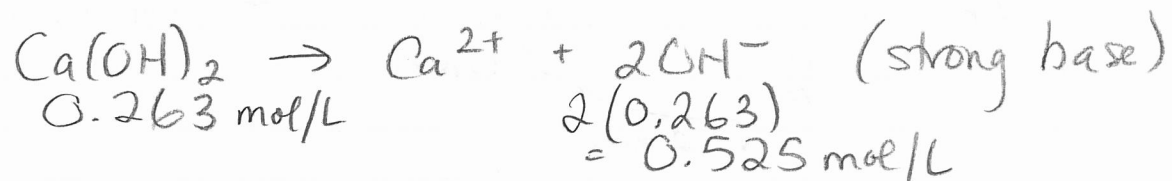
Example 3: Neutralization



$$\textcircled{3} \quad [\text{Ca}(\text{OH})_2] = \frac{n}{v} = \frac{0.0525}{0.200} = 0.263 \text{ mol/L}$$

continued on next

Solve as previous lessons



$$\begin{aligned} \text{pOH} &= -\log[\text{OH}^-] \\ &= -\log(0.525) \\ &= 0.27 \end{aligned}$$

$$\begin{aligned} \text{pH} &= 14 - \text{pOH} = 14 - 0.27 \\ &= 13.72 \end{aligned}$$

check? \uparrow
 > 7 for a base \checkmark