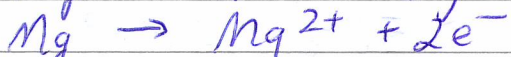
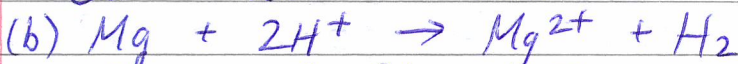
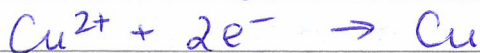
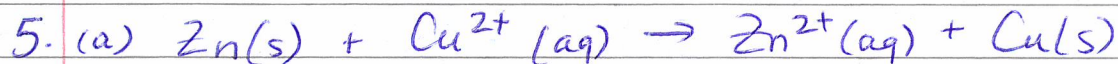
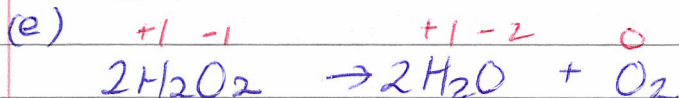
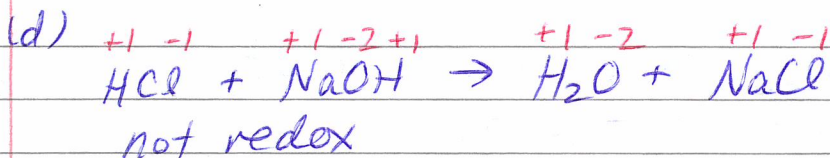
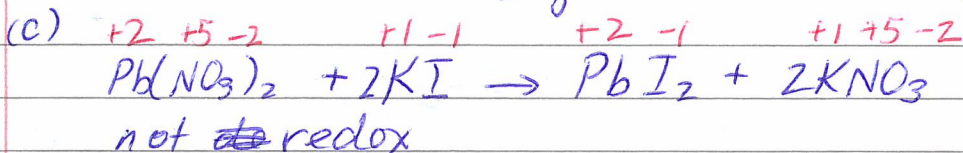
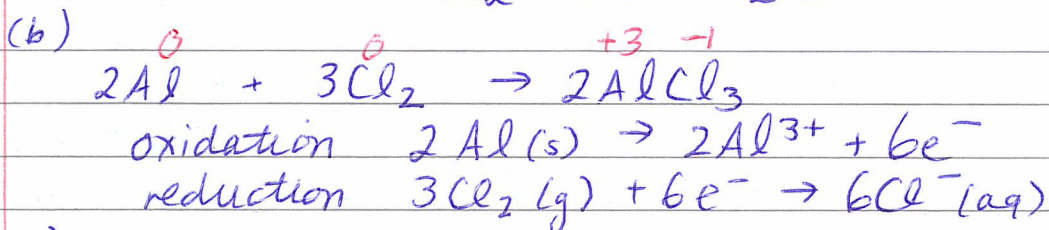
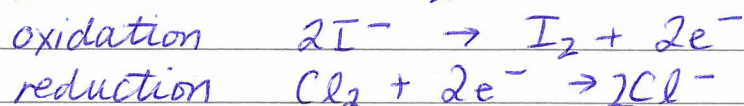
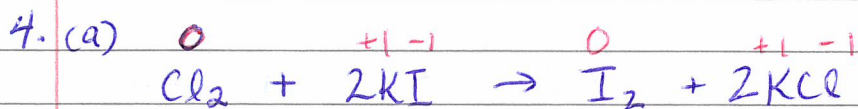
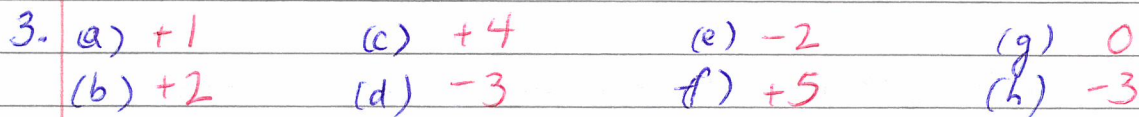
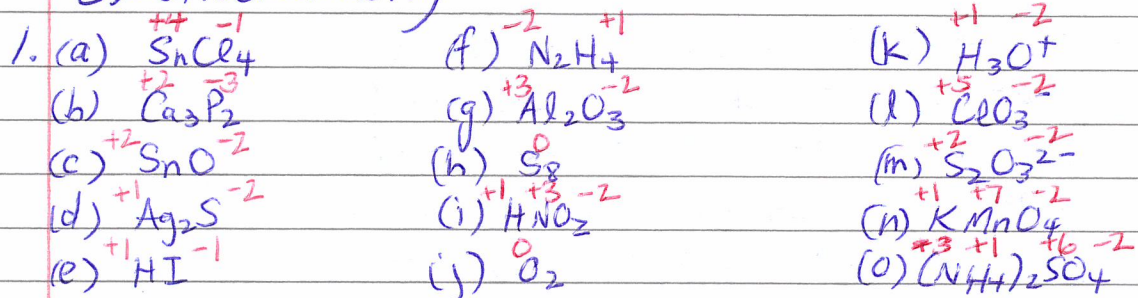
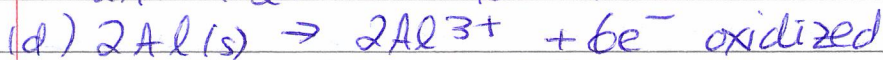
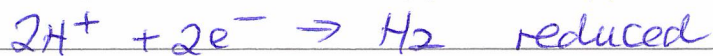
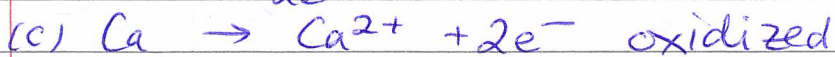
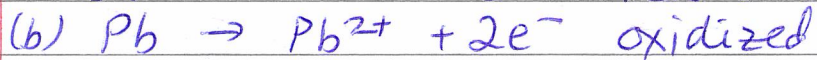
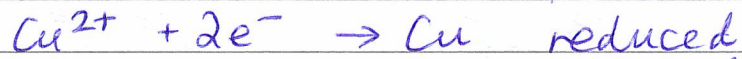
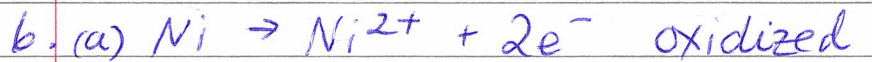
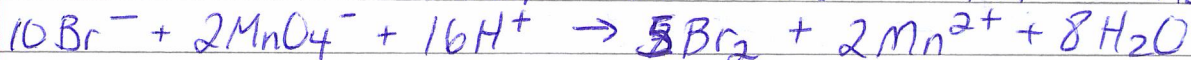
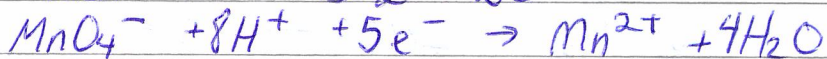
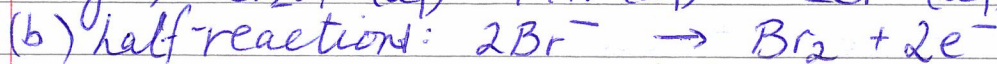
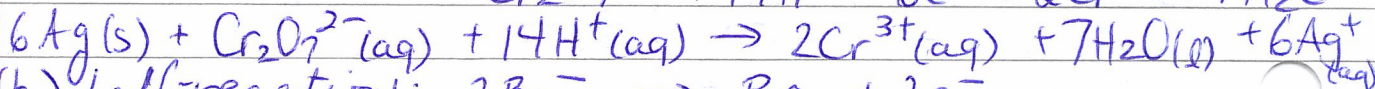
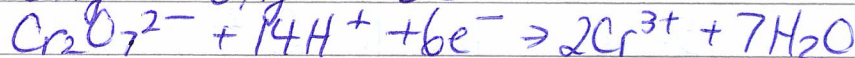
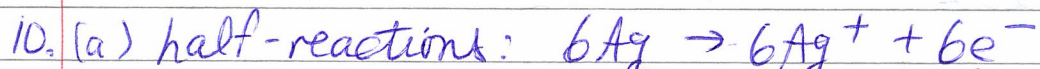
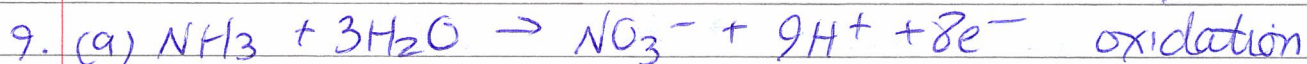
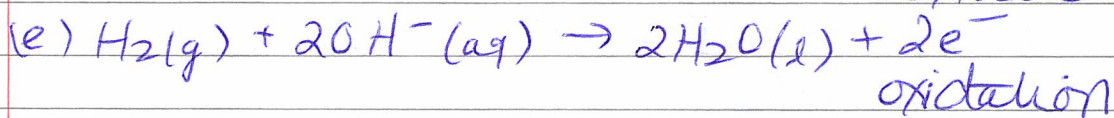
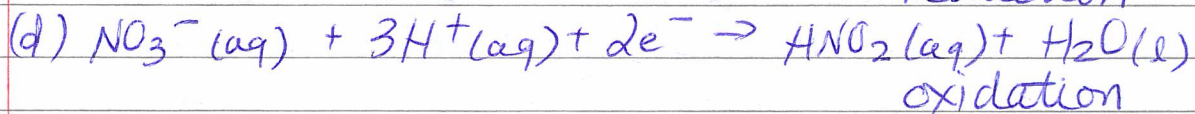
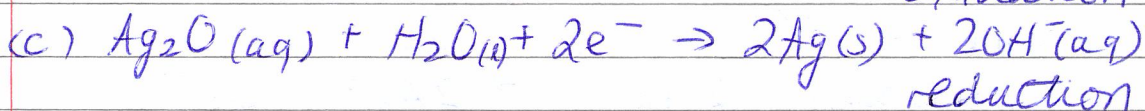
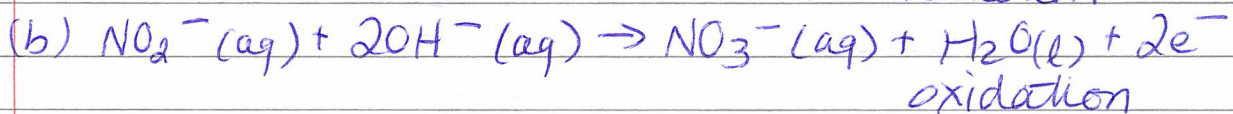
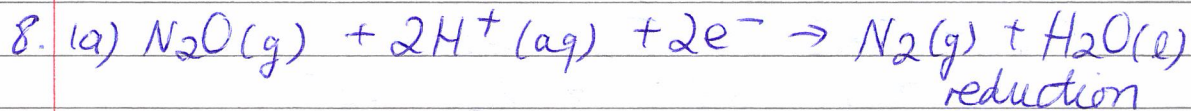


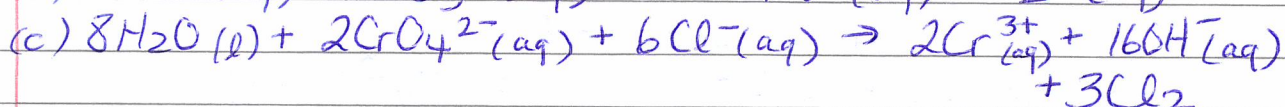
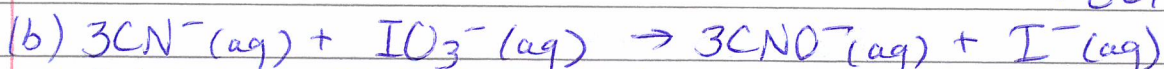
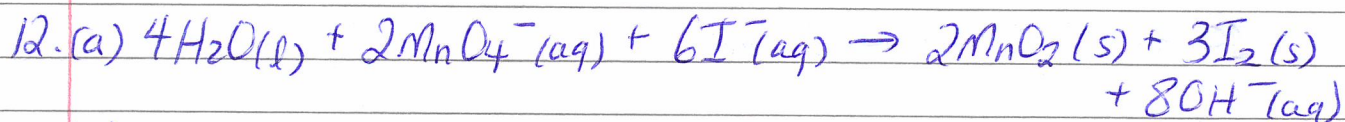
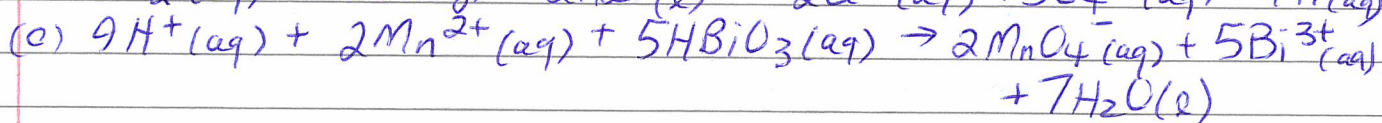
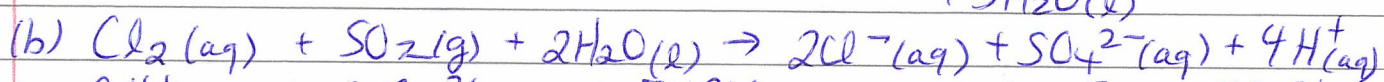
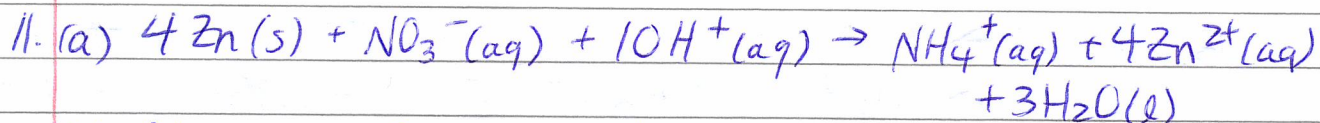
Electrochemistry





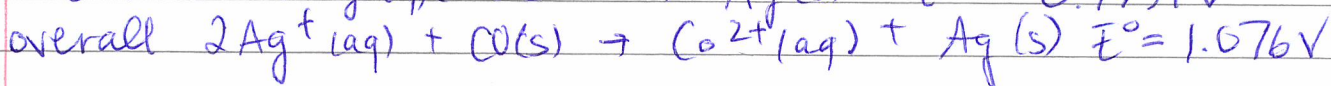
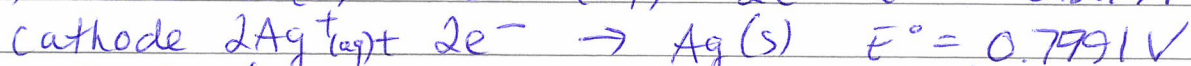
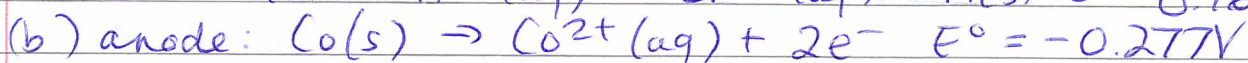
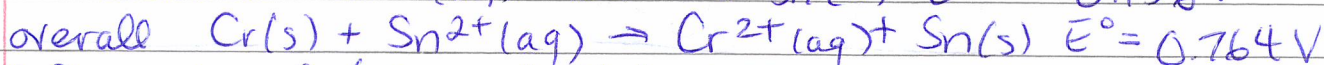
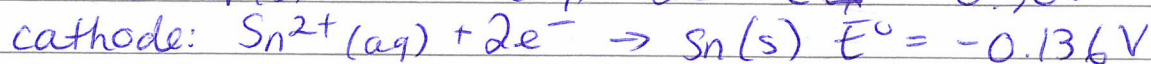
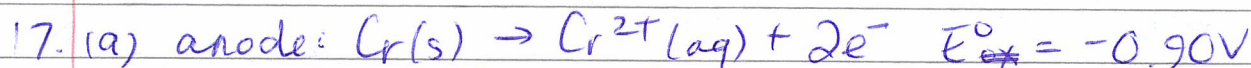
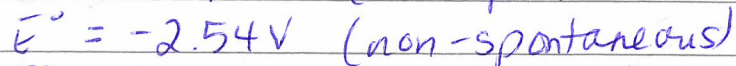
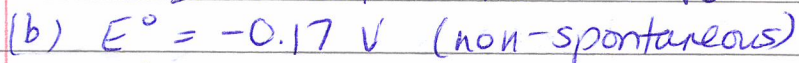
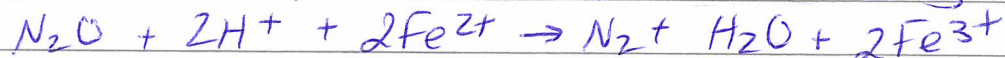
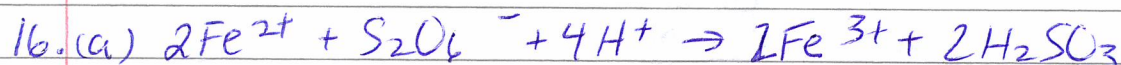
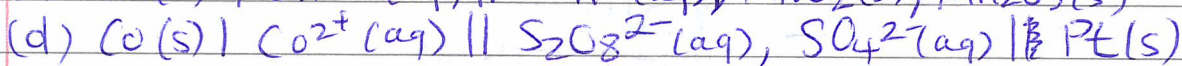
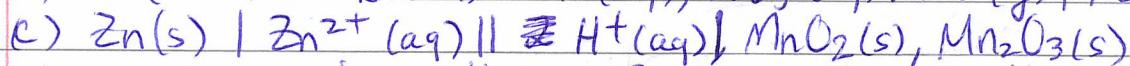
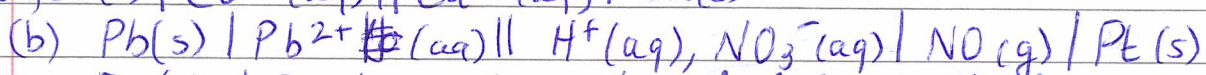
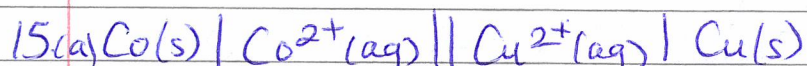
7. no - all elements have the same oxidation number as reactants and products

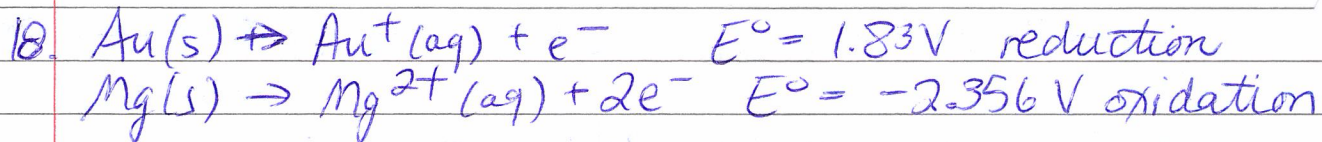




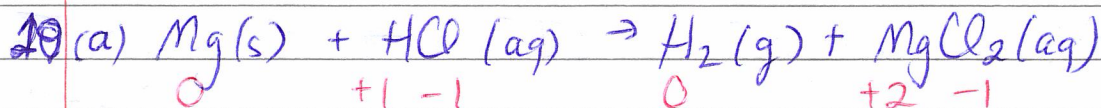
13. anode is - in galvanic and + in electrolytic

14. see chart in notes

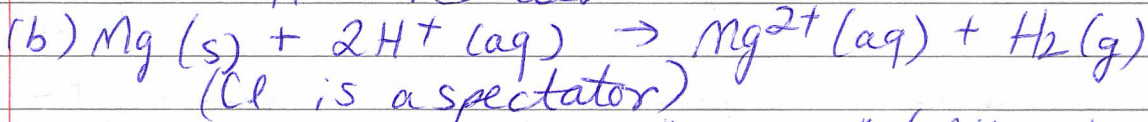




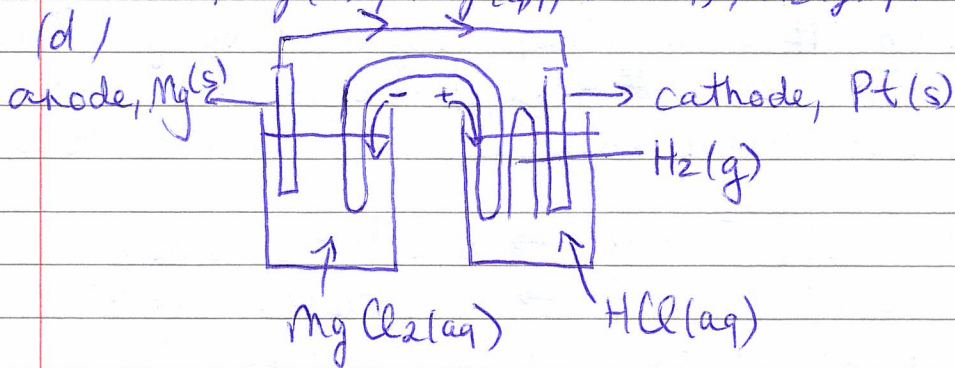
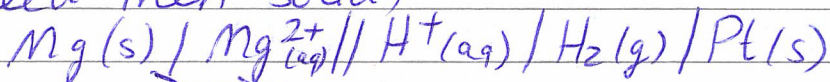
Au cathode $E^\circ = 4.186V$
 Mg anode



Mg oxidized
 H reduced



(c) Mg is anode, H_2 is "cathode" (although would need inert solid)



(e) 2.356V Spontaneous

